

상용화 추출제 및 이온성액체에 의한 염산의 용매추출

Viet Nhan Hoa Nguyen · §이만승

목포대학교 공과대학 신소재공학과

Solvent Extraction of Hydrochloric Acid Using Commercial Extractants and Synthesized Ionic Liquids

Viet Nhan Hoa Nguyen and §Man Seung Lee

Department of Advanced Materials Science & Engineering, Institute of Rare Metal, Mokpo National University, Chonnam 534-729, Republic of Korea

요 약

염산의 용매추출 및 탈거 거동을 조사하기 위해 상용화추출제인 LIX 63 단독 및 TEHA/ Cyanex 923/ Aliquat 336의 혼합추출제와 ALi-SCN과 ALi-PC 이온성액체를 사용하였다. 상기 추출제중에서 ALi-PC가 가장 높은 추출률(80%)을 보였으나 탈거가 어려웠다. 반면 LIX 63의 경우 추출률은 낮았으나, 탈거율은 81%로 가장 높았다. 옥탄올을 유기상에 첨가하면 염산 추출이 오히려 감소했다. 에탄올을 염산용액에 첨가하면 LIX 63, ALi-PC, ALi-SCN에 의한 염산의 추출과 탈거 거동이 크게 개선되었다.

주제어 : 염산화수, 용매추출, 탈거, 개질제

Abstract

The extraction and stripping of HCl from aqueous solutions by commercial extractants like LIX 63 and its mixture with TEHA/ Cyanex 923/ Aliquat 336 and ionic liquids like ALi-SCN, ALi-PC in kerosene was investigated. Among these extractants, ALi-PC showed the best extraction effectivity (above 80%), but it was difficult to strip HCl from the loaded phase. Although the extraction percentage of HCl by LIX 63 was not high, the stripping performance was above 81%. The addition of octanol to the organic phase led to negative effect on the extraction performance of HCl. The addition of C₂H₅OH into aqueous solutions significantly increased the extraction and stripping percentage of HCl by LIX 63, ALi-PC and ALi-SCN.

Key words : Hydrochloric acid recovery, Solvent extraction, Stripping, Modifier

1. Introduction

Hydrochloric acid (HCl) is widely employed in hydrometallurgical processes as a metal leaching agent

for ores and secondary resources due to its excellent chemical characteristics such as strong complex formation and fast reaction kinetics¹⁾. The chloride system applied in metallurgy can be operated over a wide range of

· Received : November 13, 2020 · Revised : December 3, 2020 · Accepted : December 10, 2020

§ Corresponding Author : Man Seung Lee (E-mail : mslee@mokpo.ac.kr)

Department of Advanced Materials Science & Engineering, Mokpo National University, 1666 Yeongsan-ro, Cheonggye-myeon, Muan-gun, Chonnam 58554, Korea

©The Korean Institute of Resources Recycling. All rights reserved. This is an open-access article distributed under the terms of the Creative Commons Attribution Non-Commercial License (<http://creativecommons.org/licenses/by-nc/3.0/>), which permits unrestricted non-commercial use, distribution and reproduction in any medium, provided the original work is properly cited.

oxidizing conditions and acidity²). In order to meet the environmental regulation against the treatment of HCl and to improve process economics, it is desirable to recover HCl from the spent leaching solutions^{1,2}.

Several methods are available to treat effluents containing inorganic acids, including neutralization, diffusion dialysis, membrane process, precipitation, solvent extraction, and ion exchange^{1,3,4}). Among these methods, solvent extraction is considered to be one of the most effective one for extracting inorganic acids from effluents. During solvent extraction, the reaction between the extractants and the protons of the acid might be responsible for the extraction of these acids⁵). In general, most tertiary amines extract hydrogen ions through the protonation mechanism, while neutral extractants can extract through the solvation⁶⁻⁹). According to previous works, the commercial extractants such as TEHA (tris 2-ethylhexyl amine), Cyanex 923 (a mixture of four trialkylphosphine oxides), TBP (tributyl phosphate), Alamine 336 (a mixture of tri-octyl/decyl amine) have been employed for the recovery of H₂SO₄ and HCl¹⁰). Among them, TEHA has the best performance for extraction but stripping percentage of HCl from the loaded organic phase was low, resulting in low concentration of HCl in the recovered acid^{6,10}). The HCl from the loaded TEHA could be completely stripped by 2 M ammonia solution¹¹). However, it seems to be an infeasible solution in industry due to the complication of stripping liquor and the difficulty in regeneration of HCl. Although the extraction percentage of H₂SO₄ and HCl by Alamine 336 is higher than that by Cyanex 923 and TBP, but the stripping of the acids from the loaded Alamine 336 is also very difficult^{6,10,12}). It has been reported that some ionic liquids (ILs) can extract hydrogen ions from weak HCl^{13,14}). The extraction of hydrogen ion by ILs depends on solution pH. The extraction reaction of hydrogen ions from weak chloride solution by ILs synthesized from Aliquat 336 (N-Methyl-N, N, N-trioctylammonium chloride) and Cyanex 272 (bis(2,2,4 trimethylpentyl)phosphinic acid) can be de-

scribed as solvation mechanism⁵).

The solvent extraction of inorganic acids from spent leaching solutions can be beneficial on the light of environment and process economics but most of the works have been done from weak acidic solutions¹²). In order to overcome some disadvantages of employing single TEHA, the influence of adding neutral extractants such as Cyanex 921 (trioctylphosphine oxide), DOS (din-octyl sulfide), MIBK (methyl isobutyl ketone), TBP (tributyl phosphate), and TOP (tri octyl phosphate) to TEHA has been investigated on the extraction of HCl¹¹). LIX 63 (7-Hydroxy-5,8-diethyl-6-dodecanone oxime) is a kind of oxime extractant and many works have been reported on the synergistic effect of the extractant mixture containing LIX 63^{15,16}). However, there are few works on the solvent extraction of HCl with the mixture of TEHA and LIX 63.

Some ionic liquids (ILs) which are synthesized by mixing Aliquat 336 and organophosphorus acids can act as bifunctional extractants and thus extract hydrogen ions. Therefore, in this work, solvent extraction of hydrochloric acid from concentrated solution was investigated by employing commercial extractants and ILs. For this purpose, the extraction and stripping behavior of HCl was compared between commercial extractants (LIX 63 and a mixture of LIX 63 and TEHA) and synthesized ILs (ALi-SCN, ALi-PC).

2. Experimental

The acidity of the solution was adjusted by adding doubly distilled water or a mixture of doubly distilled water and ethanol (C₂H₅OH; Daejung Co., 94.5%) to concentrated HCl (Daejung Co., 35%) and H₂SO₄ solution (Daejung Co., 95%). The commercial extractants such as TEHA (R₃N, BASF Co., 99%), LIX 63 (HR, Is Chem. Co., 70%), Cyanex 923 (Cytec Inc., 93%), Aliquat 336 (BASF Co., 93%) were used without any purification. 1-Octanol (Duksan Pure Chemical Co., 99%) was used as a modifier. Ionic liquid (ALi-PC) was

synthesized by mixing an equimolar concentration of Aliquat 336 and PC88A (2-ethylhexylphosphonic acid mono-2-ethylhexyl ester, HA, Cytec Inc., 95%) according to the method reported in the literature¹³. ALi-SCN was synthesized by mixing Aliquat 336 and NH₄SCN (ammonium thiocyanate, Daejung Co., 99.0%) as reported¹⁷. Kerosene (Samchun Pure Chemical Co.) was used as a diluent and all the employed chemicals were of analytical grade.

The extraction and stripping experiments were conducted by shaking equal volumes of aqueous and organic phases (20 mL each) in a screwed cap bottle for 30 mins using a Burrell wrist action shaker (model 75, USA) at room temperature (22°C ± 1). The two phases were separated using a separation funnel after holding them for 30 mins. The concentration of hydrogen ions in the aqueous phase before and after extraction was measured by volumetric titration method¹⁸. The concentration of hydrogen ions in the loaded organic phase was determined by mass balance. The experiments were repeated 3 times with an error of ± 5%.

3. Results and discussion

3.1. Solvent extraction of HCl by commercial extractants

3.1.1. Effect of nature of extractant on solvent extraction of HCl

According to the previous studies, hydrogen ions can be well extracted by amines but the stripping of hydrogen ions from the loaded amines is very difficult^{6,10,12}. Among several tertiary amines, TEHA shows the best performance for the extraction of HCl. Addition of neutral extractants such as Cyanex 921, DOS, MIBK, TBP, and TOP to TEHA has negligible effect on the extraction of HCl¹¹. Moreover, stripping of HCl from the loaded mixtures of TEHA and neutral extractants is still difficult⁶. Since stripping is the reverse reaction of extraction, it is important to find extraction systems with adequate extraction capability for HCl. It has been

reported that α-hydroxyoxime like LIX 63 can interact with HCl at high concentration by using characteristic absorption of oxime¹⁹. Therefore, a mixture of TEHA and LIX 63 was first employed in this work. In order to compare the extraction behavior between HCl and H₂SO₄, the concentration of both acids was adjusted to 3 M. Fig. 1 shows the extraction percentage of both acids by TEHA, LIX 63 and a mixture of TEHA and LIX 63. The concentration of single TEHA and LIX 63 was 50 vol% and the total volume percentage of TEHA and LIX 63 in the mixture was 50 vol% (TEHA/LIX 63 ratio = 1) and 20 vol% of octanol was added as a modifier. The extraction performance of single TEHA for both acids was similar, while no sulfuric acid was extracted by single LIX 63. Addition of TEHA to LIX 63 increased the extraction percentage of HCl from 15.7 to 25.7% and H₂SO₄ from zero to 17.6% compared to single LIX 63. This increase might be ascribed to the interaction between the extractants in binary system, which could lead to synergistic effect. The reactions can be represented as^{6,9}

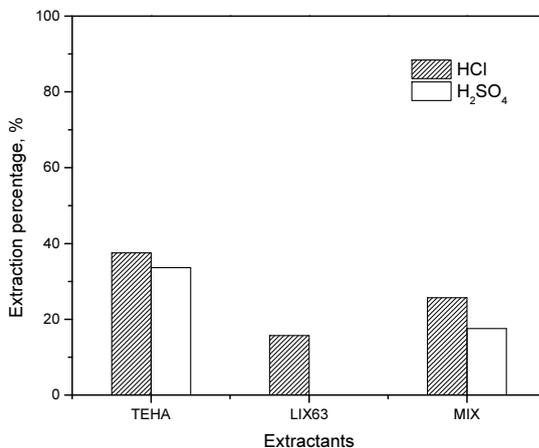
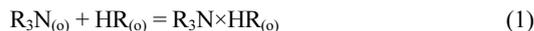


Fig. 1. Effect of extractants on hydrogen ion (3 M HCl, 3 M H₂SO₄) extraction. Condition: 50% volume of TEHA, LIX 63 and their mixture (MIX) (25:25 vol%) in kerosene and 20% octanol as modifier; O/A = 1; shaking time = 30 mins, 22°C ± 1.

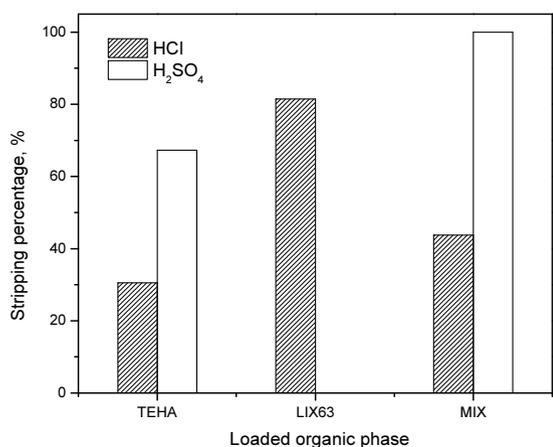
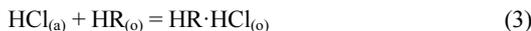
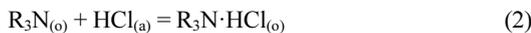


Fig. 2. Stripping of hydrogen ion from the loaded phase with H₂O at 60°C in 30 mins. Condition: organic phase: 50% volume of TEHA, LIX 63 and their mixture (MIX) (25:25 vol%) in kerosene and 20% octanol as modifier; O/A = 1.



where R₃N denotes TEHA, HR denotes LIX 63, subscript a and o represent the aqueous and organic phase, respectively.

It has been reported that temperature has a great influence on the stripping efficiency of hydrogen ions from the loaded phase by distilled water¹⁰). Therefore, the stripping of HCl and H₂SO₄ from the loaded organic phases was carried out at 60°C. The results indicate that highest stripping percentage of both acids was obtained from the loaded mixture. Especially, H₂SO₄ was completely stripped with water under from the loaded mixture of TEHA and LIX 63. Besides, HCl was also easily stripped from the organic loaded phase of LIX 63 (% stripping (%S) > 81%, A/O = 1, only one stage) (see Fig. 2).

In order to investigate the effect of the composition of the mixture, the volume percentage of LIX 63 in the mixture was varied from 10 to 40%, while that of TEHA was fixed at 50% v/v. A slight increase in extraction

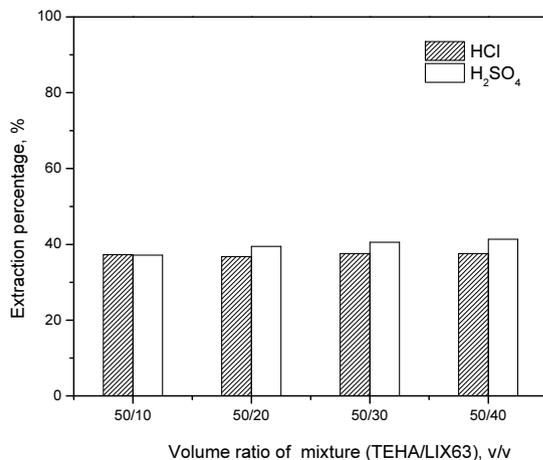


Fig. 3. Effect of concentration of mixture of TEHA and LIX 63 on hydrogen ion (HCl, H₂SO₄) extraction. Condition: extractants in kerosene and 20% octanol as modifier; O/A = 1; shaking time = 30 mins.

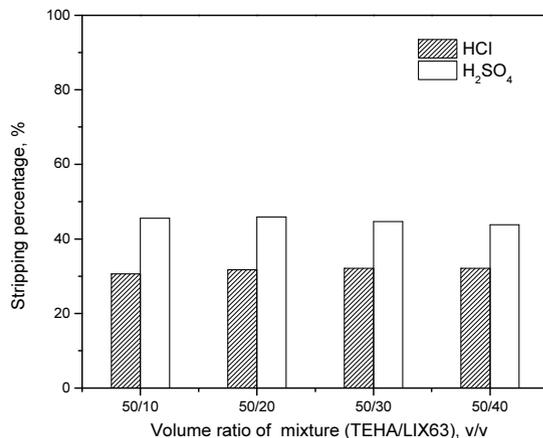


Fig. 4. Stripping of hydrogen ion from the loaded mixture TEHA and LIX 63 with H₂O at 60°C in 30 mins (O/A = 1).

percentage of both acids was observed when the volume percentage of LIX 63 increased from 10 to 40% (see Fig. 3). Namely, the extraction percentage increased from 36.8 to 37.6% and 37.1 to 41.3% for HCl and H₂SO₄, respectively. However, there was little difference in the stripping percentage of both acids in the experimental ranges (see Fig. 4). The stripping percentage of H₂SO₄ was a little higher than that of HCl with H₂O at 60°C. Fig. 4 shows that the stripping performance were less

than 50%. Compared to single extractants TEHA/ LIX 63, using their mixture slightly enhance extraction performance (Figs. 1 & 3). However, stripping of HCl from the loaded phase of mixture TEHA and LIX 63 was easier than that of single TEHA (Figs. 2 & 4). Thus, the employment of LIX 63 is more efficient for extraction of HCl on the light of the stripping step.

3.1.2. Effect of some additives on HCl extraction

In order to investigate the effect of dielectric constant of aqueous phase, C₂H₅OH was added to the aqueous phase to reduce the interaction between hydrogen ions and water molecules, which might enhance the extraction of hydrogen ions as well as hinder the extraction of water²⁰. The use of polar and protic solutes shows a favorable effect on the extraction of hydrogen ions through ion pair formation. For this purpose, HCl extraction was compared between H₂O and a mixture of H₂O and C₂H₅OH. In these experiments, the volume ratio of water and C₂H₅OH was equal. Table 1 shows the extraction percentage of HCl by LIX 63 from the mixture was 2 times higher than that from water. Although the extraction percentage of HCl was less than 10%, this result indicates that addition of C₂H₅OH to water can enhance the extraction of HCl. However, addition of too much C₂H₅OH to water can lead to the loss of extractant owing to the increase in its solubility in water mixture. Therefore, the volume ratio of H₂O and C₂H₅OH in the mixture was fixed at unity in further experiments.

Moreover, the results showed that HCl was efficiently stripped from loaded organic phases containing LIX 63. Therefore, in order to investigate the extraction and stripping behavior of HCl from the aqueous phase containing C₂H₅OH, three kinds of organic phases, namely, single LIX 63 (50%), a mixture of 50% LIX 63 and 20% Cyanex 923 and a mixture 50% LIX 63 and 20% Aliquat 336 were also employed. Among them, the mixture of LIX 63 and Aliquat 336 displays the best extraction capacity (see Table 2).

Table 1. Effect of C₂H₅OH on HCl extraction by LIX 63 50 vol% in kerosene

Aqueous phase	Extraction percentage, %
HCl in H ₂ O	4.0
HCl in H ₂ O + C ₂ H ₅ OH	8.5

*Note: aqueous phases include 1 M HCl in H₂O and 1 M HCl in the mixture of 50% H₂O and 50% C₂H₅OH (v/v), O/A = 1.

Table 2. Effect of extractant on the extraction of 1 M HCl in the mixture 50% H₂O and C₂H₅OH

Organic phase	Ratio of components, v/v	Extraction percentage, %
LIX63: kerosene	50:50	8.5
LIX63:Cyanex 923:kerosene	50:20:30	6.5
LIX63:Aliquat 336:kerosene	50:20:30	11.3

Table 3. Effect of diluent on HCl extraction by LIX 63, Aliquat 336 and their mixture

Organic phase	Ratio of components, v/v	Extraction percentage, %
LIX 63:octanol:kerosene	50:20:30	6.2
Aliquat 336:octanol:kerosene	50:20:30	0.0
LIX 63:Aliquat 336:octanol:kerosene	50:20:10:20	8.1
LIX 63:Aliquat 336:octanol:kerosene	50:20:20:10	6.2
LIX 63:Aliquat 336:octanol	50:20:30	7.1
LIX 63:Aliquat 336:kerosene	50:20:30	11.3

*Note: aqueous phases were 1 M HCl in the mixture of 50% H₂O and 50% C₂H₅OH (v/v), O/A = 1.

Modifier can affect the extraction^{4,7,21} and thus the addition of modifier like octanol to the extractants was tested. In these experiments, the aqueous was prepared by adding C₂H₅OH to water. Table 3 shows that polar modifier like octanol affects little the HCl extraction and the extraction efficiency was very low. This effect can be explained by the interaction between modifier and extractants. The interaction between extractant and modifier is proportional to the polarity of the modifier, leading to a decrease in the effective concentration of extractants²². Therefore, it can be said that addition of poplar modifier to organic phase has negative effect on HCl extraction.

3.2. Solvent extraction of HCl by synthesized ionic liquids

When metal ions are extracted by some kinds of ILs from weak acidic solutions, hydrogen ions are also extracted by the anion of the ILs and thus equilibrium pH of the aqueous phase increases^{13,14,23}. Moreover, binary mixtures consisting of Aliquat 336 and Cyanex 272 can affect the extraction of hydrogen ions from inorganic acid solutions^{13,24}. ILs which can be synthesized from Aliquat 336 and Cyanex 272 showed the excellent performance for the extraction of hydrogen ions when the initial pH range was from 0.05 to 5. These ILs can be easily prepared in the laboratory by reacting Aliquat 336 and organophosphorus acids or anion of inorganic salts through anion exchange mechanism^{5,25}.

In this work, ILs derived from Aliquat 336 like ALi-SCN and ALi-PC were employed for the extraction of HCl. Firstly, the concentration of HCl was adjusted to 1 M in aqueous phase. For comparison, water and a mixture of water and C₂H₅OH were employed. The volume percentage of H₂O and C₂H₅OH was 50% in the mixture. In these experiments, distilled water was employed as a stripping agent. As ILs, 0.5 M ALi-SCN and 0.5 M ALi-PC in kerosene were employed. Although there was little difference in the extraction percentage of HCl from water between the two ILs, the stripping percentage of HCl by ALi-SCN was much higher than that by ALi-PC (see Table 4). This may be attributed to the strong interaction of hydrogen ions with PC88A anions in the organic phase, which hindered the liberation of hydrogen ions into the aqueous phase. The addition of C₂H₅OH into water led to an increase in the extraction percentage of HCl by ALi-PC from 11.0 to 42.2%, whereas there was only a light raise from 10.2 to 12.8% in the case of ALi-SCN. However, it was noticed that stripping percentage of hydrogen ions from the loaded ALi-SCN was 90.8%. Meanwhile, hydrogen ions were not stripped from the loaded ALi-PC (see Table 5). These results can be ascribed to the change in the dielectric constant of aqueous phase by adding C₂H₅OH.

Table 4. Comparison on extraction and stripping of 1 M HCl in distilled water by ionic liquids

Extractants	Extraction percentage, %	Stripping percentage, %
ALi-SCN	10.2	22.8
ALi-PC	11.0	6.3

*Note: Organic phase: 0.5 M extractant in kerosene, O/A = 1, shaking time = 30 mins, room temperature (22°C ± 1), stripping agent: H₂O.

Table 5. Extraction and stripping of 1 M HCl in the mixture of 50% H₂O and C₂H₅OH (v/v) by ionic liquids

Extractants	Extraction percentage, %	Stripping percentage, %
ALi-SCN	12.8	90.8
ALi-PC	42.2	0.9

*Note: Organic phase: 0.5 M extractant in kerosene, O/A = 1, shaking time = 30 mins, room temperature (22°C ± 1), stripping agent: H₂O.

Besides, the hydrophobic anions of ALi-PC may stabilize the hydrogen ions in the organic phase. This means that the hydrophobic anions of ALi-PC has stronger basicity (pK_a = 4.51) than that of SCN⁻ (pK_a = 0.926) and thus these anions have stronger affinity for the hydrogen ions of HCl.

Effect of ALi-SCN concentration on the extraction of 1 M HCl from aqueous phase was also investigated. Fig. 5 shows that as ALi-SCN concentration increased from 0.3 to 1.0 M, the extraction percentage of HCl rose from 7.6 to 13.7%. Table 6 compares the effect of stripping temperature from the loaded ALi-SCN between water and the mixture of water and C₂H₅OH. Compared to the nature of the stripping agent, stripping temperature shows negligible effect. Addition of C₂H₅OH to water greatly improves the stripping percentage of HCl from the loaded ALi-SCN. Table 7 shows the same experiments from the loaded ALi-PC. Not only stripping temperature but also the nature of stripping agents did not show any effect on the stripping of HCl. The strong affinity between hydrogen ions of HCl and anion of ALi-PC might be the cause of the difficulty in stripping. Although using mixture of C₂H₅OH and water as

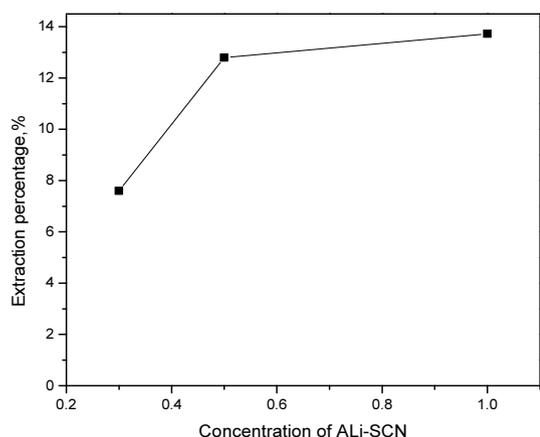


Fig. 5. Effect of ALi-SCN concentration on extraction of 1 M HCl in the mixture of 50% H₂O and C₂H₅OH. Condition: O/A = 1, shaking time = 30 mins, room temperature (22°C ± 1).

Table 6. Comparison of stripping of HCl from the loaded 1 M ALi-SCN phase

Extraction percentage, %	Stripping percentage, %		
	Stripping agent	22°C	60°C
13.7	H ₂ O	60.2	60.2
	H ₂ O + C ₂ H ₅ OH(*)	91	95.5

Note: O/A = 1, shaking time = 30 mins; aqueous phase for extraction was 1M HCl in distilled water; (): 50 vol% C₂H₅OH was added to water for stripping.

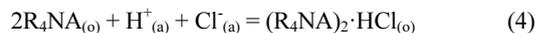
Table 7. Comparison of extraction and stripping of HCl by ALi-PC phase

Concentration of ALi-PC	Extraction percentage, %	Stripping percentage, %		
		Stripping agent	22°C	60°C
0.5 M	42.2	H ₂ O	0.8	0.8
1.0 M	80.4	H ₂ O + C ₂ H ₅ OH (*)	3.8	3.8

Note: O/A = 1, shaking time = 30 mins; aqueous phase was 1 M HCl in the mixture of 50% H₂O and 50% C₂H₅OH (v/v) for extraction; (): 50 vol% C₂H₅OH was added to water for stripping.

stripping agent did not improve stripping efficiency from the loaded phase of ALi-PC, addition of C₂H₅OH into aqueous phase was essential for enhancement of the extraction. When ALi-PC concentration increased from 0.5 to 1.0 M, the extraction percentage increased from

42.2 to 80.4% respectively (see Table 7). The solvent extraction reactions of HCl by ILs can be represented as^{5,20,26)}



where R₄NA = ALi-PC/ ALi-SCN

4. Conclusions

In order to investigate the recovery of HCl from aqueous solutions, solvent extraction experiments were performed by employing commercial extractants such as LIX 63 and its mixture with TEHA, Aliquat 336, and Cyanex 923 and synthesized ILs such as ALi-PC and ALi-SCN. Among these, ALi-PC showed higher extraction percentage than the others, but the stripping was negligible. In contrary, stripping of HCl from the loaded LIX 63 was effective (>81%). In the case of ILs, extraction performance of HCl by ALi-PC was higher than that by ALi-SCN. However, the stripping of HCl from the loaded ALi-PC was more difficult than from ALi-SCN. The extraction and stripping efficiency of HCl was enhanced by adding C₂H₅OH to the aqueous phase owing to polar modifier effect. However, addition of modifier like octanol to organic phase lowered extraction performance. Also, the effect of temperature on stripping HCl from the loaded ILs was insignificant. The difficulty in stripping of HCl from the loaded organic phase can be attributed to the excessive chemical interaction between HCl with extractant like ALi-PC, resulting in chemical conversion and formation of sustainable products in the organic phase. Further studies on the kinetics and interaction of HCl with extractants should be conducted to verify the interaction.

Acknowledgments

This work was supported by the Technology Innovation Program (Development of Material Component Technology) (Project number: 20013071) funded by the Ministry of Trade, Industry & Energy (MOTIE, Korea).

References

- McKinley, C., and Ghahreman, A., 2017 : Hydrochloric acid regeneration in hydrometallurgical processes: a review, *Mineral Processing and Extractive Metallurgy*, 127(3), pp.157-168.
- Van Weert, G., and Peek, E. M. L., 1992 : Reagent recovery in chloride hydrometallurgy - some missing links, *Hydrometallurgy*, 29(1-3), pp.513-526.
- Haghshenas, D. F., Darvishi, D., Rafieipour, H., et al., 2009 : A comparison between TEHA and Cyanex 923 on the separation and the recovery of sulfuric acid from aqueous solutions, *Hydrometallurgy*, 97(3-4), pp.173-179.
- Shin, C. H., Kim, J. Y., Kim, J. Y., et al., 2009 : Recovery of nitric acid from waste etching solution using solvent extraction, *J. Hazard. Mater.*, 163(2-3), pp.729-734.
- Le, M. N., Son, S. H., and Lee, M. S., 2019 : Extraction Behavior of Hydrogen Ion by an Ionic Liquid Mixture of Aliquat 336 and Cyanex 272 in Chloride Solution, *Korean J. Met. Mater.*, 57(3), pp.162-169.
- Banda, R., Nguyen, T. H. and Lee, M. S., 2013 : Recovery of HCl from chloride leach solution of spent HDS catalyst by solvent extraction, *Chem Process Eng-Inz.*, 34(1), pp.153-163.
- Stas, J., and Alsawaf, H., 2016 : Liquid - liquid extraction of hydrochloric acid from aqueous solutions by Tri-n-dodecylamine and Tri-n-octylamine/diluents, *Period. Polytech. Chem. Eng.*, 60(2), pp.130-135.
- Vieux, A. S. and Rutagengwa, N., 1976 : Organic phase species in the extraction of hydrochloric acid by tri-n-octylamine in various organic diluents, *J Phys Chem.*, 80(12), pp.1283-1291.
- Alguacil, F. J., and López, F. A., 1996 : The extraction of mineral acids by the phosphine oxide Cyanex 923, *Hydrometallurgy*, 42(2), pp.245-255.
- Kesieme, U., Chrysanthou, A., Catulli, M. et al., 2018 : A review of acid recovery from acidic mining waste solutions using solvent extraction, *J. Chem. Technol. Biotechnol.* 93(12), pp.3374-3385.
- Le, M. N., Nguyen, T. H. and Lee, M. S., 2018 : Extraction and stripping behavior of hydrochloric acid from aqueous solution by Cyanex 923/TEHA and its mixtures, *Geosystem Eng.*, 22(3), pp.1-9.
- Sarangi, K., Padhan, E., Sarma, P. V. R. B., et al., 2006 : Removal/recovery of hydrochloric acid using Alamine 336, Aliquat 336, TBP and Cyanex 923, *Hydrometallurgy*, 84(3-4), pp.125-129.
- Fortuny, A., Coll, M. T. and Sastre, A. M., 2012 : Use of methyltrioctyl/decylammonium bis 2,4,4-(trimethylpentyl) phosphinate ionic liquid (ALiCY IL) on the boron extraction in chloride media, *Sep. Purif. Technol.*, 97, pp.137-141.
- Castillo, J., Coll, M. T., Fortuny, A., et al., 2014 : Cu(II) extraction using quaternary ammonium and quaternary phosphonium based ionic liquid, *Hydrometallurgy*, 141, pp.89-96.
- Truong, H., Lee, M. and Son, S., 2017: Extraction of Palladium(II) from Hydrochloric Acid Solutions by Solvent Extraction with Mixtures Containing Either Cyanex 301 or LIX 63, *Metals*, 7(12), pp.541.
- Wang, L. Y., and Lee, M. S., 2017 : Synergistic extraction of Co(II) over Ni(II) from chloride solutions by a mixture of Cyanex 301 and LIX 63, *Geosystem Eng.*, 20(6), pp.311-317.
- Preston, J. S., 1982 : Solvent Extraction of Cobalt(II) and Nickel(II) by a Quaternary Ammonium Thiocyanate, *Sep Sci Technol.*, 17(15), pp.1697-1718.
- Vogel A. I., 1989 : *Textbook of Quantitative Chemical Analysis*, Longman Scientific & Technical, New York, 5th ed., pp.295-296.
- Sato, T., Sato, K. and Noguchi, Y., 2001 : Liquid-liquid extraction of iron (III) from hydrochloric acid solutions by α -hydroxyoxime, *Shigen-to-Sozai*, 117(3), pp.226-229.
- Eyal, A. M. and Canari, R., 1995 : pH Dependence of Carboxylic and Mineral Acid Extraction by Amine-Based Extractants: Effects of pKa, Amine Basicity, and Diluent Properties, *Ind. Eng. Chem. Res.*, 34(5), pp.1789-1798.
- Eyal, A. M., Hazan, B. and Bloch, R., 1991 : Recovery and concentration of strong mineral acids from dilute solutions through LLX. II. Reversible extraction with branched-chain amines, *Solvent Extr. Ion Exch.*, 9(2), pp.211-222.
- Liu, Y. and Lee, M. S., 2016 : Effect of the diluents on the interaction between components in the binary mixtures of organophosphorus acid and tertiary amine, *J. Mol. Liq.*, 220, pp.41-48.
- Rout, A., Venkatesan, K. A., Srinivasan, T. G. et al., 2012 : Ionic liquid extractants in molecular diluents: Extraction behavior of europium (III) in quaternary ammonium-based ionic liquids, *Sep. Purif. Technol.*, 95, pp.26-31.
- Liu, Y., Jeon, H. S. and Lee, M. S., 2015 : Extraction of hydrochloric acid with binary mixtures of tertiary amine

- and organophosphorus acid and analysis of the interaction between the constituents of these mixtures, Hydrometallurgy, 155, pp.44-50.
25. Dupont, D., Depuydt, D. and Binnemans, K., 2015 : Overview of the Effect of Salts on Biphasic Ionic Liquid/Water Solvent Extraction Systems: Anion Exchange, Mutual Solubility, and Thermomorphic Properties, J. Phys. Chem. B, 119(22), pp.6747-6757.
26. Tran, T. T. and Lee, M. S., 2020 : Interactions between ionic liquid (ALiCY) and TBP and their use for extraction of Co(II) and Ni(II) in hydrometallurgy, Korean J. Met. Mater., 58(6), pp.1-10.

Viet Nhan Hoa Nguyen

- CAN THO 대학교 석사
- 현재 목포대학교 신소재공학과 박사과정

이만승

- 현재 목포대학교 신소재공학과 교수
 - 당 학회지 제11권 1호 참조
-