

DFT Study of Bis(Crown-Ether) Analogue of Tröger's Base Complexed with Bisammonium Ions: Hydrogen Bonds

Kwangho Kim and Jong-In Choe*

Department of Chemistry, Chung-Ang University, Seoul 156-756, Korea. *E-mail: choeji@cau.ac.kr

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The optimized structures and complexation energies of bis(18-crown-6-ether) analogue (**2**) of Tröger's base (**1**) with a series of primary alkylbisammonium ions have been calculated by *DFT* B3LYP/6-31G(d,p) method. The calculated complexation efficiency (−142.84 kcal/mol) of **2** for butane-1,4-diylbisammonium guest is better than twice of the value (−61.40 kcal/mol) for butylammonium ion. The multiple hydrogen-bond abilities for the complexes are described as the function of the length of the alkyl substituents of the bisammonium guests with *normal*-alkyl chain $[-(\text{CH}_2)_n-]$, $n = 4-8$. The longer bisammonium guest shows the stronger hydrogen-bonding characterizations (the distance and the quasi-linear angle of the $\text{N-H}\cdots\text{O}$) to the host **2** than the shorter bisammonium ions. These calculated results agree with the experimental data of the complexation of **2** with bisammonium salts ($[\text{NI}_3(\text{C}_n\text{H}_{2n})\text{NH}_3] \text{Cl}_2$).

Key Words : *DFT* B3LYP/6-31G(d,p), Bis(crown-ether), Bisammonium ion, Complex

Introduction

Molecular recognition of organic ammonium guests is very important in view of many functional biogenic ammonium ions, and a number of sophisticated host systems are derived from crown ethers and calixarenes.^{1,2} Selective *endo*-calix complexation of alkylammonium ions by functionalized (1,3)-*p*-*tert*-butylcalix[5]crown ether³ and by calix[5]arene-based molecular vessels⁴ have been reported. Finding new receptors⁵ for organic ammonium salts are also important in the perspective that bioactive molecules such as amino acids, peptides and proteins exist partly as ammonium ions in aqueous solutions and for the enantioselective separation of small chiral primary ammonium salts.⁶ It is well established that primary ammonium ions are effective complexation partners for 18-crown-6 derivatives.⁷ The bonding forces constitute a combination of charge-dipole and hydrogen-bonding interactions.⁸ Quite a few existing receptors for primary bisammonium ions are based on bis-crown ethers.⁹

Previously, we have calculated the structures of *p*-*tert*-butylcalix[4]crown-6-ether and its alkylammonium complexes by using a HF/6-31G method.¹⁰ We have also reported the *ab initio* calculation results for the energies and structures of *p*-*tert*-butylcalix[6](aza)cryptand and its alkylammonium complexes.^{11a}

The first objective of this research is to calculate the relative binding affinity of various alkylbisammonium ions with bis(18-crown-6-ether) analogue (**2**) of Tröger's base (**1**). Another objective is to compare multiple hydrogen-bonding characteristics for the $[2+(\text{NH}_3(\text{CH}_2)_n\text{NH}_3)]^{2+}$ complexes as the function of the length of the primary alkyl substituents of the bisammonium guests by using *DFT* B3LYP/6-31G(d,p)¹⁴ method which includes the effect of an electron correlation and the basis set with a polarization function.

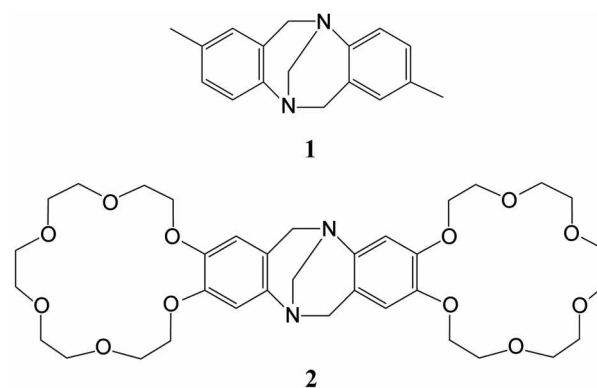


Figure 1. Chemical Drawings of (a) Tröger's base **1**, (b) bis(18-crown-6-ether) analogue **2** of the Tröger's base.

Computational Methods

The initial structures of host **2** and bisammonium ions were created by HyperChem.¹³ In order to find optimized conformations, we executed conformational search by simulated annealing method described in previous research.¹⁶ *DFT* B3LYP/6-31G(d,p) full optimization of the structure of host **2** or its bisammonium complex took more than several days to reach the final structure for each complex by using Gaussian 98¹⁷ on supercomputer (NEC SX-7) at the Okazaki National Research Institute of Japan. Normal mode frequencies of the final structures also have been calculated. Each vibrational spectrum shows no negative value of frequency, which suggests that the optimized structure exists in the minimum point.

Results and Discussion

The B3LYP/6-31G(d,p) complexation energies of the host **2** with the primary bisammonium ions are summarized in

Table 1. B3LYP/6-31G(d,p) Binding Energies (kcal/mol)^a of Host **2**, Alkyl Bisammonium Cations and their Complexes

Host 2 +Guest Complexation	Guest ^b					
	<i>n</i> -Butyl _m	<i>n</i> -Butyl	<i>n</i> -Pentyl	<i>n</i> -Hexyl	<i>n</i> -Heptyl	<i>n</i> -Octyl
ΔE (kcal/mol) ^c	-61.40	-142.84	-135.56	-134.46	-129.74	-112.98
$\Delta\Delta E$ (kcal/mol) ^d		0.00	7.28	8.38	13.10	29.86

^aError limits in these calculations are about 0.01 kcal/mol. ^b*n*-Butyl_m = *n*-butylammonium, *n*-Butyl = *n*-butane-1,4-diylbisammonium, *n*-Pentyl = *n*-pentane-1,5-diylbisammonium, *n*-Hexyl = *n*-hexane-1,6-diylbisammonium, *n*-Heptyl = *n*-heptane-1,7-diylbisammonium, *n*-Octyl = *n*-octane-1,7-diylbisammonium ion. ^cComplexation energy (ΔE) = $E_{\text{Complex}} - E_{\text{Host}} - E_{\text{Guest}}$. ^d $\Delta\Delta E = \Delta E - \Delta E$ (*n*-Butyl) : between the bisammonium complexes.

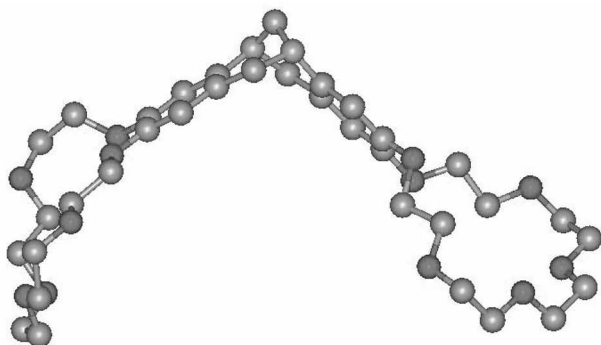
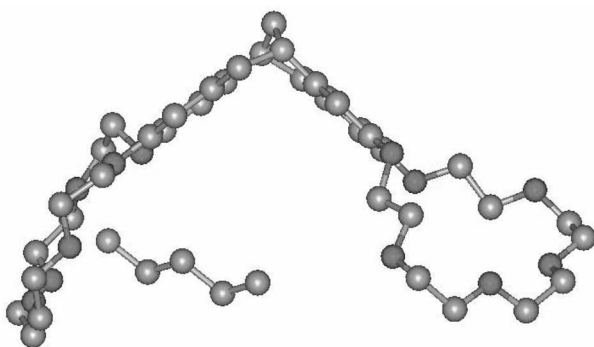
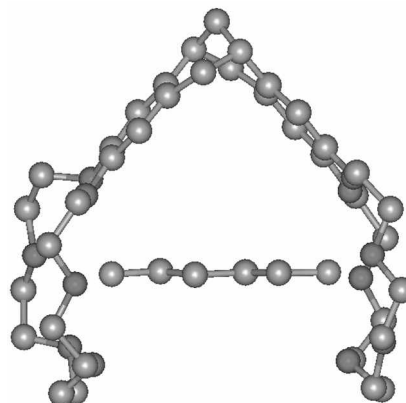
**Figure 2.** The calculated structure of free host **2**. Figure 2 is drawn without hydrogen atoms for clarity.

Table 1. An interesting comparison of the bis- and the mono-butylammonium ion tells us that the calculated complexation efficiency (-142.84 kcal/mol) of **2** for the butane-1,4-diylbisammonium guest is better than twice of the value (-61.40 kcal/mol) for the butylammonium ion. The trend of the complexation efficiencies for **2** using the *DFT* method is in the order of butane-1,4 > pentane-1,5 ~ hexane-1,6 > heptane-1,7 > octane-1,8-diylbisammonium ion.

The *DFT* optimized structure for the host **2** is described in Figure 2 which created by POSMOL¹⁸ is drawn without hydrogen atoms for clarity. The free host **2** shows the distorted structure of crown-6-ether moieties.

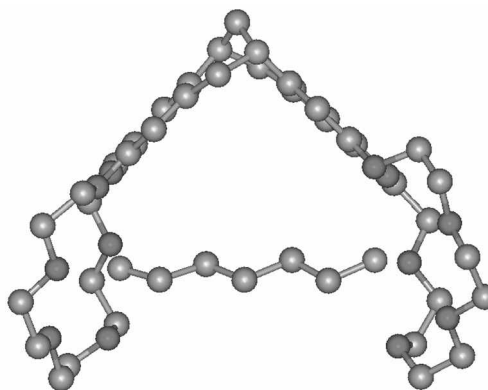
Figure 3 displays the optimized structure of the **2**+ $[\text{CH}_3(\text{CH}_2)_3\text{NH}_3^+]$ complex. Figures 4 through 8 show the *DFT* optimized complexes of **2** with various primary bisammonium ions. The shape of the title compound (**2**) looks more C_{2v} -symmetric upon complexation with a bisammonium ion, due to multiple hydrogen bonds between both of

**Figure 3.** The structure of the $[\mathbf{2}+(n\text{-butylammonium})]^+$ complex ($\mathbf{2}+\text{CH}_3(\text{CH}_2)_3\text{NH}_3^+$).**Figure 4.** The structure of $[\mathbf{2}+(\text{butane-1,4-diylbisammonium})]^{2-}$ complex ($\mathbf{2}+\text{NH}_3^+(\text{CH}_2)_4\text{NH}_3^-$).

the crown-6-ether moieties of the host and the double ends of the bisammonium cation. These six hydrogen bonds provide the remarkable stability of the $\mathbf{2}+\text{R}(\text{NH}_3^+)_2$ complex. When one observes the Figure 4, two crown-6-ether rings in the $[\mathbf{2}+(\text{butane-1,4-diylbisammonium})]^{2-}$ complex become more vertical due to the multiple hydrogen bonds. As the length of the primary alkyl chain of the bisammonium guest gets longer, two 18-crown-6-ether moieties of host **2** become further apart.

Table 2 shows the optimized geometrical parameters (N \cdots O and H \cdots O distances, N-H bond lengths and N-H \cdots O angles) of the bisammonium complexes to understand the strength of the hydrogen bonding.¹⁹

Table 3 summarizes the individual and the average hydrogen-bond distances from nitrogen atoms of the bisammonium

**Figure 5.** The structure of $[\mathbf{2}+(\text{pentane-1,5-diylbisammonium})]^{2-}$ complex ($\mathbf{2}+\text{NH}_3^+(\text{CH}_2)_5\text{NH}_3^-$).

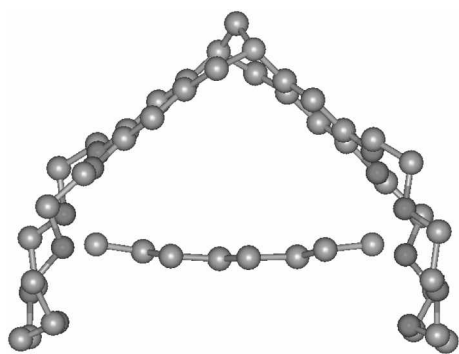


Figure 6. The structure of [2+(hexane-1,6-diylbisammonium)]²⁻ complex (2+NH₃⁺(CH₂)₆NH₃⁺).

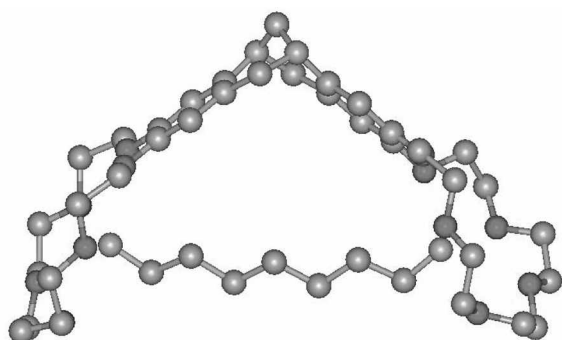


Figure 7. The structure of [2+(heptane-1,7-diylbisammonium)]²⁻ complex (2+NH₃⁺(CH₂)₇NH₃⁺).

nium ion to the oxygen atoms of host **2**. These H-bond distances [$d(\text{N}\cdots\text{O}) \sim 2.91 \text{ \AA}$] are comparable with the experimental values of pentane-1,5-diylbisammonium (2.88 Å) and hexane-1,6-diylbisammonium (2.94 Å) complexes of tetra-*p*-sulfonylcalix[4]arene.²⁰ However, our calculated H-bond distance (2.91 Å) display shorter, therefore, stronger than the values calculated from other primary monoammonium complexes of calix[4]crown-6-ether¹⁰ (~3.05 Å) and calix[6](aza)cryptand^{11a} (~3.15 Å), and also shorter than the

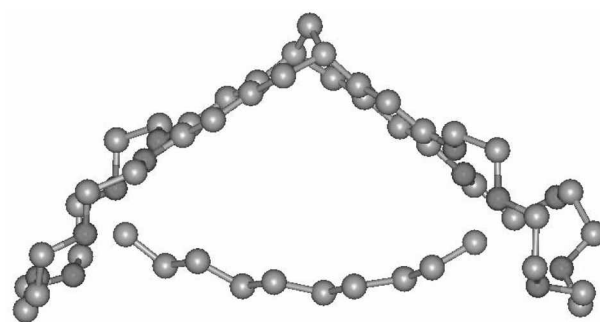


Figure 8. The structure of [2+(octane-1,8-diylbisammonium)]²⁻ complex (2+NH₃⁺(CH₂)₈NH₃⁺).

Table 2. Geometrical parameters (N \cdots O and H \cdots O distances, N-H bond lengths and N-H \cdots O angles — in Å and °) of Various Bisammonium Complexes of **2**

Guest ^a	N \cdots O	H \cdots O	N-H	N-H \cdots O
NH ₃ ⁺ (CH ₂) ₄ NH ₃ ⁺	2.911	1.886	1.037	170.822
NH ₃ ⁺ (CH ₂) ₅ NH ₃ ⁺	2.936	1.916	1.036	168.612
NH ₃ ⁺ (CH ₂) ₆ NH ₃ ⁺	2.920	1.897	1.036	170.700
NH ₃ ⁺ (CH ₂) ₇ NH ₃ ⁺	2.904	1.872	1.035	174.549
NH ₃ ⁺ (CH ₂) ₈ NH ₃ ⁺	2.897	1.888	1.034	166.733

^aSee the footnote b of Table 1.

experimental values of calix[6](aza)cryptand^{11b} (2.97-3.11 Å).

The multiple hydrogen-bonding abilities for the complexes are described as the function of the length of the alkyl substituents of the bisammonium guests. The complex with the longer alkyl (heptyl or octyl) chain of the bisammonium ion shows the shorter H-bond distances, therefore, the stronger hydrogen-bondings to **2** than the pentane-1,5-diylbisammonium ion. Although the conditions of calculation (in vacuum) and experiment (in MeOH-d₄:CDCl₃ (1:1)) are different, our *DFT* calculated results agree with the experimental data of the complexation of **2** with bis-

Table 3. Individual and Average Distances (Å) from the Nitrogen Atoms of Bisammonium Ions to the Oxygen atoms of Host **2**

Guest	Nitrogen to Ether Oxygen						Average
	N1-O1	N1-O2	N1-O3	N2-O4	N2-O5	N2-O6	
NH ₃ ⁺ (CH ₂) ₄ NH ₃ ⁺	2.914	2.888	2.908	2.907	2.946	2.902	2.911
NH ₃ ⁺ (CH ₂) ₅ NH ₃ ⁺	3.014	2.939	2.870	2.920	2.930	2.942	2.936
NH ₃ ⁺ (CH ₂) ₆ NH ₃ ⁺	2.951	2.928	2.882	2.951	2.928	2.882	2.920
NH ₃ ⁺ (CH ₂) ₇ NH ₃ ⁺	2.961	2.910	2.851	2.880	2.941	2.879	2.904
NH ₃ ⁺ (CH ₂) ₈ NH ₃ ⁺	2.939	2.860	2.962	2.844	2.930	2.849	2.897

Table 4. Individual and Average Angles (°) of the N-H \cdots O atoms of the Bisammonium Complexes of **2**

Guest	Nitrogen-H \cdots Oxygen						Average
	N1-H \cdots O1	N1-H \cdots O2	N1-H \cdots O3	N2-H \cdots O4	N2-H \cdots O5	N2-H \cdots O6	
NH ₃ ⁺ (CH ₂) ₄ NH ₃ ⁺	167.30	168.29	178.79	172.33	164.66	173.56	170.82
NH ₃ ⁺ (CH ₂) ₅ NH ₃ ⁺	171.11	152.53	170.24	174.56	171.52	171.71	168.61
NH ₃ ⁺ (CH ₂) ₆ NH ₃ ⁺	171.14	163.72	177.23	171.13	163.73	177.24	170.70
NH ₃ ⁺ (CH ₂) ₇ NH ₃ ⁺	176.72	178.63	176.98	169.87	173.39	171.72	174.55
NH ₃ ⁺ (CH ₂) ₈ NH ₃ ⁺	173.90	160.69	174.66	168.96	170.98	151.21	166.73

ammonium salts ($[\text{NH}_3(\text{CH}_2)_n\text{NH}_3] \text{Cl}_2$).¹²

Table 4 lists the individual and average angles of the N-H...O atoms of the bisammonium complexes of **2**. The quasi-linear angles with the average range of 166-175° provide the remarkable strengths of the hydrogen bonds for the bisammonium complexes. These calculated distances and angles of multiple N-H...O hydrogen bonds in the primary bisammonium complexes of **2** may provide an excellent prototype of receptor for bisammonium ammonium salts.

Conclusion

By using *DFT* B3LYP/6-31G(d,p) methods we have optimized the structures and have calculated the complexation energies of the bis(18-crown-6-ether) analogue (**2**) of Tröger's base (**1**) with various primary bisammonium ions. The complex with the longer alkyl chain (heptyl or octyl) of the bisammonium ion shows the shorter N-H...O distance, therefore, stronger hydrogen bond than the pentane-1,5-diylbisammonium ion. The average hydrogen-bonded distance [$d(\text{N}\cdots\text{O}) = \sim 2.91 \text{ \AA}$] of five different alkylbisammonium complexes displays shorter than those of the other alkylmonoammonium complexes. The quasi-linear angles of the N-H...O atoms of the bisammonium complexes with the average range of 166-175° provide the remarkable strengths of the hydrogen bonds for the bisammonium complexes.

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